

ODYSSEY Molecular Explorer

— Release 7.0 —

Correlation with the

Massachusetts Science Learning Standards Chemistry, High School

Updated October 2006

Content Standards

1. Properties of Matter

Central Concept: Physical and chemical properties reflect the nature of the interactions between molecules or atoms, and can be used to classify and describe matter.

1.1 Identify and explain physical properties (e.g., density, melting point, boiling point, conductivity, malleability) and chemical properties (e.g., the ability to form new substances). Distinguish between chemical and physical changes.

→ **C12** *Chemical Matter* "Types of Properties"

→ **C18** *Chemical Matter* "Identifying by Density"

1.2 Explain the difference between pure substances (elements and compounds) and mixtures. Differentiate between heterogeneous and homogeneous mixtures.

→ **C3** *Chemical Matter* "Examples of Elements"

→ **C4** *Chemical Matter* "Types of Compounds"

→ **C5** *Chemical Matter* "Types of Mixtures"

1.3 Describe the three normal states of matter (solid, liquid, gas) in terms of energy, particle motion, and phase transitions.

→ **C6** *Chemical Matter* "States of Matter"

→ **C7** *Chemical Matter* "Comparing States Side-by-Side"

→ **C13** *Chemical Matter* "Physical Changes"

→ **H9** *Liquids & Solids* "Molecular Motion and Physical States"

→ **H20** *Liquids & Solids* "Melting Transition"

2. Atomic Structure and Nuclear Chemistry

Central Concepts: Atomic models are used to explain atoms and help us understand the interaction of elements and compounds observed on a macroscopic scale. Nuclear chemistry deals with radioactivity, nuclear processes, and nuclear properties. Nuclear reactions produce tremendous amounts of energy and lead to the formation of elements.

2.2 Describe Rutherford's "gold foil" experiment that led to the discovery of the nuclear atom. Identify the major components (protons, neutrons, and electrons) of the nuclear atom and explain how they interact. .

→ **D2** *Atoms* "Distribution of Mass in Atoms"

2.4 Write the electron configurations for the first twenty elements of the periodic table. .

→ **D4** *Atoms* "Hydrogen Atom"

→ **D8** *Atoms* "Atomic Orbitals" .

→ **D9** *Atoms* "Comparing Helium, Neon, and Argon"

→ **D14** *Atoms* "Orbitals of a Krypton Atom"

3. Periodicity

Central Concepts: Repeating (periodic) patterns of physical and chemical properties occur among elements that define families with similar properties. The periodic table displays the repeating patterns, which are related to the atoms' outermost electrons.

3.1 Explain the relationship of an element's position on the periodic table to its atomic number. Identify families (groups) and periods on the periodic table.

→ **C3** *Chemical Matter* "Examples of Elements"

→ **P1** *Main Groups & Transition Metals* "Alkali Metals"

→ **P2** *Main Groups & Transition Metals* "Alkaline Earth Metals"

→ **P3** *Main Groups & Transition Metals* "Boron Group"

→ **P4** *Main Groups & Transition Metals* "Carbon Group"

→ **P6** *Main Groups & Transition Metals* "Nitrogen Group"

→ **P7** *Main Groups & Transition Metals* "Oxygen Group"

→ **P10** *Main Groups & Transition Metals* "Halogens"

→ **P11** *Main Groups & Transition Metals* "Noble Gases"

→ **P12** *Main Groups & Transition Metals* "Elements of the d- and f-Blocks"

4. Chemical Bonding

Central Concept: Atoms bond with each other by transferring or sharing valence electrons to form compounds.

4.1 Explain how atoms combine to form compounds through both ionic and covalent bonding. Predict chemical formulas based on the number of valence electrons.

→ **F1** *Chemical Bonding* "The Attraction Between Ions"

→ **F7** *Chemical Bonding* "Electron Sharing"

→ **F8** *Chemical Bonding* "Energetics of Covalent Bonding"

4.2 Draw Lewis dot structures for simple molecules and ionic compounds.

→ *Many Stockroom Pages*

4.3 Use electronegativity to explain the difference between polar and nonpolar covalent bonds.

→ **F11** *Chemical Bonding* "Polar Bonds and Molecules"

→ **F13** *Chemical Bonding* "Classifying by Bond Polarity"

4.4 Use valence-shell electron-pair repulsion theory (VSEPR) to predict the molecular geometry (linear, trigonal planar, and tetrahedral) of simple molecules.

→ **F14** *Chemical Bonding* "VSEPR Theory"

→ **F15** *Chemical Bonding* "Comparing Shapes"

4.5 Identify how hydrogen bonding in water affects a variety of physical, chemical, and biological phenomena (e.g., surface tension, capillary action, density, boiling point).

→ **H19** *Liquids & Solids* "Liquid Water"

4.6 Name and write the chemical formulas for simple ionic and molecular compounds, including those that contain the polyatomic ions: ammonium, carbonate, hydroxide, nitrate, phosphate, and sulfate.

→ **C20** *Chemical Matter* "Naming Compounds"

→ **F10** *Chemical Bonding* "Polyatomic Ions"

5. Chemical Reactions and Stoichiometry

Central Concepts: In a chemical reaction, one or more reactants are transformed into one or more new products. Chemical equations represent the reaction and must be balanced. The conservation of atoms in a chemical reaction leads to the ability to calculate the amount of products formed and reactants used (stoichiometry).

5.1 Balance chemical equations by applying the laws of conservation of mass and constant composition (definite proportions).

→ **M3** *Kinetics* "Mechanism of a Reaction"

6. States of Matter, Kinetic Molecular Theory, and Thermochemistry

Central Concepts: Gas particles move independently of each other and are far apart. The behavior of gas particles can be modeled by the kinetic molecular theory. In liquids and solids, unlike gases, particles are close to each other. The driving forces of chemical reactions are energy and entropy. The reorganization of atoms in chemical reactions results in the release or absorption of heat energy.

6.1 Using the kinetic molecular theory, explain the behavior of gases and the relationship between pressure and volume (Boyle's law), volume and temperature (Charles's law), pressure and temperature (Gay-Lussac's law), and the number of particles in a gas sample (Avogadro's hypothesis). Use the combined gas law to determine changes in pressure, volume, and temperature.

- **G13** *Gases* "Pressure-Volume Relationship"
- **G14** *Gases* "Boyle's Law"
- **G16** *Gases* "Pressure and Temperature"
- **G18** *Gases* "Avogadro's Law"
- **G19** *Gases* "Universality of the Ideal Gas Law"

6.2 Perform calculations using the ideal gas law. Understand the molar volume at 273 K and 1 atmosphere (STP).

- **G9** *Gases* "Standard Conditions"
- **G19** *Gases* "Universality of the Ideal Gas Law"

6.3 Using the kinetic molecular theory, describe and contrast the properties of gases, liquids, and solids. Explain, at the molecular level, the behavior of matter as it undergoes phase transitions.

- **C6** *Chemical Matter* "States of Matter"
- **C7** *Chemical Matter* "Comparing States Side-by-Side"
- **C13** *Chemical Matter* "Physical Changes"
- **H9** *Liquids & Solids* "Molecular Motion and Physical States"
- **H20** *Liquids & Solids* "Melting Transition"

6.4 Describe the law of conservation of energy. Explain the difference between an endothermic process and an exothermic process.

- **L4** *Thermochemistry* "Vibrating Diatomic Molecule"
- **M2** *Kinetics* "Reactive Collisions"
- **M3** *Kinetics* "Mechanism of a Reaction"
- **N2** *Equilibria* "Equilibrium and Temperature"

6.5 Recognize that there is a natural tendency for systems to move in a direction of disorder or randomness (entropy).

- **O1** *Chemical Thermodynamics* "Gas Expansions"
- **O3** *Chemical Thermodynamics* "Heat Conduction"

7. Solutions, Rates of Reaction, and Equilibrium

Central Concepts: Solids, liquids, and gases dissolve to form solutions. Rates of reaction and chemical equilibrium are dynamic processes that are significant in many systems (e.g., biological, ecological, geological).

7.1 Describe the process by which solutes dissolve in solvents.

→ **I2** *Solutions* "Process of Dissolving"

→ **I11** *Solutions* "Energetics of Solution Formation"

7.2 Calculate concentration in terms of molarity. Use molarity to perform solution dilution and solution stoichiometry.

→ **I6** *Solutions* "Concentration of a Dissolved Pesticide"

→ **I7** *Solutions* "Molarity vs. Molality"

7.5 Identify the factors that affect the rate of a chemical reaction (temperature, mixing, concentration, particle size, surface area, catalyst).

→ **M2** *Kinetics* "Reactive Collisions"

7.6 Predict the shift in equilibrium when a system is subjected to a stress (LeChatelier's principle) and identify the factors that can cause a shift in equilibrium (concentration, pressure, volume, temperature).

→ **N2** *Equilibria* "Equilibrium and Temperature"

→ **N3** *Equilibria* "Equilibrium and Pressure"

8. Acids and Bases and Oxidation-Reduction Reactions

Central Concepts: Acids and bases are important in numerous chemical processes that occur around us, from industrial procedures to biological ones, from the laboratory to the environment. Oxidation-reduction reactions occur when one substance transfers electrons to another substance, and constitute a major class of chemical reactions.

8.1 Define the Arrhenius theory of acids and bases in terms of the presence of hydronium and hydroxide ions in water and the Bronsted-Lowry theory of acids and bases in terms of proton donors and acceptors.

→ **K1** *Acids & Bases* "Strong Acids"

8.2 Relate hydrogen ion concentrations to the pH scale and to acidic, basic, and neutral solutions. Compare and contrast the strengths of various common acids and bases (e.g., vinegar, baking soda, soap, citrus juice).

→ **K2** *Acids & Bases* "Comparing Oxoacids"

Mathematical Skills

Students are expected to know the content of the *Massachusetts Mathematics Curriculum Framework*, through grade 8. Below are some specific skills from the *Mathematics Framework* that students in this course should have the opportunity to apply:

- ✓ Construct and use tables and graphs to interpret data sets.

→ *Many Labs*

- ✓ Solve simple algebraic expressions.

→ *Many Labs*

- ✓ Convert within a unit (e.g., centimeters to meters).

→ *Many Labs*

- ✓ Use common prefixes such as *milli-*, *centi-*, and *kilo-*.

→ *Many Labs*

The following skills are not detailed in the *Mathematics Framework*, but are necessary for a solid understanding in this course:

- ✓ Determine the correct number of significant figures.

→ **G6** *Gases* "Gas Pressure"

→ **I6** *Solutions* "Concentration of a Dissolved Pesticide"

- ✓ Use the Celsius and Kelvin scales.

→ **G8** *Gases* "Temperature Scales"